

KRASIL'NIKOV, Gennadiy Aleksandrovich; SOLENOK, Z.A., inzh.,
retsenzent; SOKOL'SKIY, I.F., red.; USTINOVA, V.A.,
tekhn. red.

[Oil and water cooling of the transformers of the V.I.
Lenin Volga Hydroelectric Power Station] Masliano-
vodianoe okhlazhdeniye transformatorov Volzhskoi TES im.
V.I.Lenina. Moskva, Gos.energ.izd-vo, 1960. 46 p.
(MIRA 16:10)

(Electric transformers--Cooling)
(Volga Hydroelectric Power Station (Lenin))

KRASIL'NIKOV, G.V.

Complete inertialess pneumatic-electric automation of the
section for the heat treatment of raw materials. Spirt.
prom. 22 no.4:27-29 '56. (MLRA 10:2)

1. Lomovskiy spirtovyy zavod.
(Automatic control) (Alcohol)

KRASIL'NIKOV, G.V.

"Instruments and regulators for the alcohol industry" by V.V. Aronovich. Reviewed by G.V. Krasil'nikov. Spirt. prom. 23 no.5:43-44 '57.
(Distilling industries--Equipment and supplies) (NLPA 10:8)
(Aronovich, V.V.)

LARIN, M.N., prof., doktor tekhn.nauk; KRASIL'NIKOV, I.M.; TSYGANOVA, M.P.; AKIMOV, A.V., kand.tekhn.nauk; BUDNIKOV, N.Ye., inzh.; PETROSYAN, L.K., kand.tekhn.nauk; DIBNER, L.G., inzh.; SILAYEVA, I.D., inzh.; MAGAZINER, Z.G., kand.tekhn.nauk; UVAROVA, A.F., tekhn.red.

[Cutting tools designed for high production and their efficient operation] Vysokoproizvoditel'nye konstruksii reztsov i ikh ratsional'naya ekspluatatsiya. Pod red. M.N.Larina. Moskva, Gos.nauchno-tekhn.izd-vo mashinostroit.lit-ry, 1959. 239 p. (MIRA 12:6)

1. Moscow. Vsesoyuznyy nauchno-issledovatel'skiy instrumental'nyy institut. 2. Sotrudniki Vsesoyuznogo nauchno-issledovatel'skogo instrumental'nogo instituta (for all except Uvarova).

(Metal-cutting tools)

KRASIL'NIKOV I.F., inzh.; TEKIMISHCHYAN, A.V., kand. tekhn. nauk

Selection of a plan for cooling hydraulic coupling for the
drive of a powerful centrifuge. Obog. i brik. ugl. no.26:
25-32 '62. (MIRA 17.8)

SAYAPIN, YU. I., Eng.; ERASIL'NIKOV, I. YE., Eng.;
KAD'KOV. V. P., Eng.

Electric Power

Electric energy consumption in the preparation of reinforcements for hydrotechnical concrete. Hidr. stroi. 21 no. 7, 1952.

Monthly List of Russian Accessions, Library of Congress, December 1952. UNCLASSIFIED.

KRASIL'NIKOV, I.Ye.

Improving the operation of NIIPOL dryers. Obn. tekhn. opyt. [PULP]
no. 4:8 '56. (MIRA 11:10)
(Drying apparatus--Textile fabrics)

111 AND 110 ORDERS

PROCESSES AND PROPERTIES INDEX

BC

System $\text{CaO-SiO}_2\text{-H}_2\text{O}$. Sorption of calcium oxide by silica gel. K. G. Krasilevich and A. V. Kisilev (*J. Phys. Chem. Russ.*, 1944, 18, 327-328).—The amount of CaO taken up from aq. Ca(OH)_2 by SiO_2 gel within, say, 1 hr. increases with $[\text{Ca(OH)}_2]$ (0-45 m-equiv. per l.) according to the usual adsorption isotherm. 1 month later the sorption isotherm has a vertical branch at 2.2 m-equiv. of Ca(OH)_2 per l.; the composition of the solid phase is $\text{CaO} \cdot \text{SiO}_2 \cdot \text{aq.}$ and the solution contains 6-7 m-equiv. of SiO_2 per l. At higher $[\text{Ca(OH)}_2]$ no equilibrium is reached even within 17 months; probably more basic silicates are formed. Two native forms of hydrated SiO_2 behave like SiO_2 gel. J. J. B.

COMMON ELEMENTS

MATERIAL INDEX

ASB-114 METALLURGICAL LITERATURE CLASSIFICATION

111 AND 110 ORDERS

111 AND 110 ORDERS

1ST AND 2ND SERIES										3RD AND 4TH SERIES									
COMMON ELEMENTS										COMMON VARIANTS									
CA										2									
<p>specific area and heat of wetting of asbestos fiber. A. V. Kiselev and K. O. Kravtchuk (Moscow Textile Inst.). <i>J. Applied Chem. (U.S.S.R.)</i> 19, 316-31(1946).-- The specific area of dry type-3 asbestos fiber, measured by adsorption of butyl alc. or butyric acid from CCl₄ soln., is 2.10×10^4 sq. cm./g., as compared with an external fiber surface area of 400 sq. cm./g. measured by the micro- scope. The integral molar heat of wetting of dry, out- gassed (by evacuation) material is 10.3 ± 0.1 kg.-cal./mol. Cyrus Feldman</p>																			
ASS. S.A. METALLURGICAL LITERATURE CLASSIFICATION										EDWIN BOWLING									
EDWIN DIVISION										EDWIN DIVISION									
EDWIN DIVISION										EDWIN DIVISION									

DZHIOT, O. M., KISELEV, A. V. and KRASIL'NIKOV, K. G.

"Ca pillary," DOK. AN, 58, No. 3, 1947

CLASSIFICATION		PROCESSES AND PROPERTIES INDEX		2	
C A		<p>Adsorption Isotherms from solutions on layering or crystallization. Adsorption of phenol on silica gel from solutions in heptane. K. G. Kravtchikov and A. V. Kisilev. Doklady Akad. Nauk S.S.S.R., 63, 603 (1948). On a fine pore (I) and a coarse-pore (II) silica gel, characterized by a max. of the effective pore radii distribution curve at, resp., 15-20 Å. and 50 Å., and by liquid H₂O adsorption at satn. of, resp., 0.63 and 0.92 cc./g., adsorption curves of PhOH in CCl₄ soln. at 20° are similar and show no S-shape. At 40°, i.e. at the temp. of layering, the adsorption curve on II becomes S-shaped, whereas the curve on I remains normal. At the completion of capillary layering, at high relative concns., PhOH and CCl₄ are adsorbed in a ratio corresponding to the compn. of the PhOH-rich layer. The limiting amts. of PhOH adsorbed at 40°, estd. from the isotherms, are, on I and II, resp., 4.2 and 7.4 millimoles/g., corresponding to 0.50 and 0.90 cc./g., resp. These vols. are close to the vols. of liquid H₂O adsorbed at satn., which fact indicates dense filling of the pores at completion of capillary layering, in contrast to the case of crystn. at 20°. The absence of an S-shaped rise on I indicates that its pores are practically filled already at low relative concns.; on II, the early stage of adsorption results only in coverage of the pore walls, the pore space being filled only at the stage of capillary layering.</p> <p>N. Thon</p>			
MOSCOW State Univ. im. Lomonosov					
ASH-SLA METALLURGICAL LITERATURE CLASSIFICATION		ECONOMY NOMINATIONS			
ECONOMY SYMBOLISM		ECONOMY ONE ORV LIST			
SYMBOLS WITH ONLY ONE		BIBLIOGRAPHY			
A B C D E F G H I J K L M N O P Q R S T U V W X Y Z		A B C D E F G H I J K L M N O P Q R S T U V W X Y Z			

KRASIL'NIKOV, K. G. i KISHLEV, A. V.

26220 Priroda sorbtsii CaO iz vodnykh rastverov silikagelyani i aluminogelyani
(Sistemy CaO---SiO₂-H₂O i CaO-Al₂O₃-H₂O) Sbornik nauch. rabot po vyaznushchim
materialam. M., 1949, s. 141-52^{3 2}

SO: LETOPIS' NO. 35, 1949

KRASIL'NIKOV, K. G.

"Adsorption from Solution by porous Bodies near the Critical Mixing Temperature.
The System Silica Gel-Acetic Acid - Heptane," Dokl. AN SSSR, 69, No. 6, 1949.

Inst. Physical Chemistry, Dept. Chem. Sci., Acad. Sci. USSR

KRASIL'NIKOV, K. G.

"Effect of the Structure of the Silica Gel on the Velocity of the Sorption of Calcium Hydroxide from Aqueous Solutions," O. M. Dzhigit, A. V. Kiselev, and K. G. Krasil'nikov (Gosudarst. Vsesoyuz. Nauch-Issledovatel. Inst. Tsement. Prom. and Moskov. Gosudarst. Univ. im M. V. Lomonosova). Dokl. Akad. Nauk SSSR, 71, 77-9, (1950) - The amts. of Ca(OH)_2 , in mg. -equiv./g., sorbed from a clear aq. soln. after a stated length of time (1 hr to 30 days), are plotted against the concn. of the soln. after sorption. The isotherms are substantially different for a coarsely-porous silica gel (I), characterized by marked capillary condensation and considerable hysteresis in the sorption of C_5H_{12} vapor at 20° , and a finely-porous silica gel (II) showing no capillary condensation under the same conditions. Pore vol. distribution curves show, for II, a sharp peak at about 10 A., and in the range of 80-100 A. for I. Sorption of Ca(OH)_2 was detd. with fractions of I and II remaining after sifting with 10,000 mesh/sq. cm., and heated 4 hrs. at 350° . All points of the isotherms corresponding to the same initial concn. lie on the same straight line which connects the point on the axis of abscissas expressing the original concn. of the soln. In the case of I, the 1-hr. isotherm shows irregularities of shape indicative of vol. sorption. Isotherms taken at later stages become increasingly straightened out; the 24-hr. isotherm is very nearly vertical, and, after 30 days, it corresponds to the equil. between the initial silica gel $\text{SiO}_2\text{aq.}$, the silicate $\text{CaO.SiO}_2\text{aq.}$, and the aq. soln. The coarse pores of I permit ready diffusion of Ca(OH)_2 , and the Ca silicate formed does not prevent its access to the surface of yet unreacted SiO_2 . This is not so in the case of II. All isotherms, including that taken after 30 days show the familiar shape of initial rise and leveling off, and lie very closely one above the other. Sorption after 30 days is only a little greater than after 1 day.

2

CA

Unimolecular and multimolecular adsorption from solution on a nonporous adsorbent. A. V. Kiselev and K. G. Krasil'nikov (Moscow State Univ.). *Doklady Akad. Nauk S.S.S.R.* 77, 1831-4 (1951).—Adsorption isotherms from solns. of PhOH in C_6H_6 on $BaSO_4$ powder of sp. surface area 0.7 sq. m./g. were detd. at 20°, where satn. leads to crysln. of PhOH, and at 40° where satn. results in sepn. into 2 liquid layers. At 20°, the isotherm levels off at a relatively low concn., 25 millimole/l. (i.e. 0.1 of satn.) and the adsorption remains practically unchanged up to satn. of the soln. and beyond it. This const. adsorption, 58 millimoles/g., corresponds to a surface area of 28 A.²/mol. PhOH; consequently, under these conditions, only one close-packed unimol. adsorption layer is formed. At 40°, the isotherm is S shaped, beginning to rise at approx. a relative concn. $c/c_s = 0.6$ ($c_s = 1.3$ mole/l.). From this point on, adsorption is multimol., but the thickness of the adsorbed layer does not exceed 2-3 mols. even at $c/c_s = 0.9$. A similar isotherm is found with MeOH in C_6H_6 on $BaSO_4$. N. Thon

1951

KRASIL'NIKOV, K.G.
C. 2

1951

General and Physical Chemistry

Absolute adsorption isotherms from solution. K. G. Krasil'nikov and A. V. Kiselev (Moscow State Univ.) *Doklady Akad. Nauk S.S.S.R.* 77, 1047-48 (1951); cf. *Argul, et al., C.A.* 45, 8464g. Adsorption isotherms were detd. for solns. of PhOH in $C_{12}H_{26}$ at 20° and 40°, on the highly coarse silica gel "E" (I), the coarse silica gel "VK-Kh-2" (II), a relatively fine-pore silica gel "No. 45" (III), and nonporous $BaSO_4$ (IV). Total pore vol. in cc./g., most frequent pore diam. in Å, and sp. surface area S in sq. m./g. were: for I 1.73, 200, 320; II 0.92, 90, 350; III 0.53, 30, 600; IV —, —, 9.7. At 20°, ads. in the vol. results in crystn., and at 40° in sepa. into 2 liquid layers. The isotherm on I at 20° has the normal shape up to satn., whereas at 40° it is S-shaped. Plots of the amt. of PhOH adsorbed, a , per unit sp. surface area, S , as a function of its concn., c , at 20°, coincide for I, II, and III; consequently, the porosity of the silica gel shows no effect on the adsorption up to a pore size of 30 Å. Near satn., a/S = micromole/sq. m., which gives a/S = 26 Å²/mol. PhOH, i.e. adsorption at 20° remains unimol. up to satn. Conversely, adsorption of PhOH from soln. in $C_{12}H_{26}$ can be used to det. the unknown sp. surface area of a silica gel. On IV, adsorption is still unimol., but the limit is reached at about $c = 0.1$ c.m. Plots of a/S at 40°, on I, II, and IV, are still S-shaped, with the isotherms coinciding only in their initial portions but diverging beyond the initial stage. This divergence is the result of multimol. adsorption on IV, but is due to capillary layering in the case of the silica gels; the latter effect is more pronounced, the larger is the pore vol. That the rise of a/S on coarse silica gels at higher concn. is not due to mere multimol. adsorption, or to layering in the space between grains, follows from the fact that in the range of relative concns. 0.7-1.0, a/S is much greater, and rises more steeply with c , for the silica gels than for IV. Consequently, the effect consists in capillary layering, analogous to capillary condensation of vapors. At $c/c_s = 1$, I adsorbs 16-17 millimoles PhOH per g., or 1.5 cc. liquid PhOH/g. This, with some $C_{12}H_{26}$, is sufficient for complete filling of the pore vol. N. Thon

YEGOROV, M.M.; KISELEV, V.F.; KRASIL'NIKOV, K.G.; SIMANOV, Yu.P.

Effect of the phase composition of adsorbents in the system
 $\text{Al}_2\text{O}_3 - \text{H}_2\text{O}$ on their surface activity. *Izv.vys.ucheb.zav.;*
khim.i khim.tekh. 2 no.3:360-365 '59. (MIRA 13:8)

1. Moskovskiy gosudarstvennyy universitet im. M.V.Lomonosova,
kafedra fiziki. (Surface chemistry)
(Aluminum oxide)

KRASIL'NIKOV, K. G.

KRASIL'NIKOV, K. G. -- "Investigation of Sorption of Hydrated Calcium Oxide by Silica." Sub 12 May 52, Moscow Order of Lenin Chemicotechnological Inst imeni D. I. Mendeleev. (Dissertation for the Degree of Candidate in Technical Sciences).

SO: Vechernaya Moskva, January-December 1952

KISELEV, A. V., KRASIL'NIKOV, K. G., POKROVSKIY, N. L., AVGUL', N. N., DZHIGIT, O. M.,
SHCHERBAKVA, K. D.

Heat of Wetting

Dependence of the heat of wetting of silica gel with water on the surface coverage. Zhur.
fiz. khim. 26 no. 7, 1952.

Monthly List of Russian Accessions. Library of Congress, December 1952. Unclassified.

KRASIL'NIKOV, K. G.

234T20

and in heat of adsorption in the region of concns close to the point of sepn (layer formation) apparently are related to polymol adsorption. At that point on the adsorption isotherm, where the curve has a stepwise character, the heat of adsorption curve has a series of minima. Presented by Acad M. M. Dubinin 5 Jul 52.

234T20

"Dok Ak Nauk SSSR" Vol 86, No 1, pp 111-113

Sci USSR

"Studying the Adsorption and Heat of Adsorption of Phenol in Aqueous Solution on Nonporous Carbon Black," V. F. Kiselev, K. G. Krasil'nikov, Moscow State U

Acad of Sci USSR

USSR/Chemistry - Adsorption

1 Sep 52

KRASIL'NIKOV, K. G.

The effect of dehydration of a silica-gel surface on its adsorption properties. A. V. Kiselev, K. G. Krasil'nikov, and L. M. Boboleva. *Dokl. Akad. Nauk S.S.S.R.* 141, 88-9 (1964). The effect of the chem. structure and the state of an adsorbent surface on the ads. values of the adsorption and the energy of adsorption was detd. on partially dehydrated SiO_2 gel. The adsorbates were MeOH and C_6H_6 vapors and phenol in γ -heptane soln. The removal of OH groups from the SiO_2 -gel surface by dehydration lowers the adsorption of all of the adsorbates but it is particularly significant for those that can form H bonds with the OH groups of the adsorbent surface (MeOH , $\text{C}_6\text{H}_5\text{OH}$). J. Rovtar Leach

2. Inst. Physical Chemistry, ACAD SCI USSR

KRASIL'NIKOV, K. G.

USSR /Chemical Technology. Chemical Products
and Their Application

I-12

Silicates. Glass. Ceramics. Binders.

Abs Jour: Referat Zhur - Khimiya, No 9, 1957, 31640

Author : Krasil'nikov K. G.

Title : Chemical Processes in Dispersed Bodies (Study
of the System $\text{CaO} - \text{SiO}_2 - \text{H}_2\text{O}$).

Orig Pub: Tr. Soveshchaniya po khimii tsementa. M.,
Promstroyizdat, 1956, 351-380

Abstract: Description of the results of the study of equilibrium in the $\text{CaO-SiO}_2\text{-H}_2\text{O}$ system, at different values of the CaO/SiO_2 ratio, conducted by utilizing modern adsorption research methods and, in individual cases, of electron-microscopic,

Card 1/2

USSR /Chemical Technology. Chemical Products
and Their Application

I-12

Silicates. Glass. Ceramics. Binders.

Abs Jour: Referat Zhur - Khimiya, No 9, 1957, 31640

petrographic, roentgenographic and adsorptional
structure method of analysis.

Card 2/2

YEGOROV, M.M.; KRASIL'NIKOV, K.G.; SYSOYEV, Ye.A.

Water wetting heats of various silica gels with reference to their degree of hydration. Dokl.AN SSSR 108 no.1:103-106 My '56.

(MLRA 9:8)

1. Moskovskiy gosudarstvennyy universitet imeni M.V. Lomonosova.
Predstavleno akademikom M.M. Dubininym.

(Silica) (Heat of wetting)

KRASIL'NIKOV, K. G.

USSR/ Physical Chemistry - Surface phenomena. Adsorption. Chromatography.
Ion exchange

B-13

Abs Jour : Referat Zhur - Khimiya, No 4, 1957, 11393

Author : Aleksandrova G.I., Kiselev V.F., Krasil'nikov K.G., Murina V.V.,
Sysoyev Ye.A.

Inst : Academy of Sciences USSR

Title : Heat of Wetting of Silicagel of Different Degrees of Hydration by
Some Organic Liquids

Orig Pub : Dokl. AN SSSR, 1956, 108, No 2, 283-286

Abstract : Determined were the heat values of wetting of surface unit of dehydrated, at 300-900°, of specimens of silicagel (SG) of different porosity by absolute methanol (Q_1), n-propanol (Q_2) and non-polar n-heptane (Q_3). Q_1 does not depend on the nature of porosity of SG; Q_2 and Q_3 are higher in the case of coarsely porous SG, than for finely porous, which is attributed to the effect of pores which increases on transition to larger molecules of C_3H_7OH and C_7H_{14} . Q_1 and Q_2 increase linearly with degree of hydration (θ H_2O) of SG surface, which confirms (see reference) the assum-

1/2

USSR/ Physical Chemistry - Surface phenomena. Adsorption, Chromatography.
Ion exchange

B-13

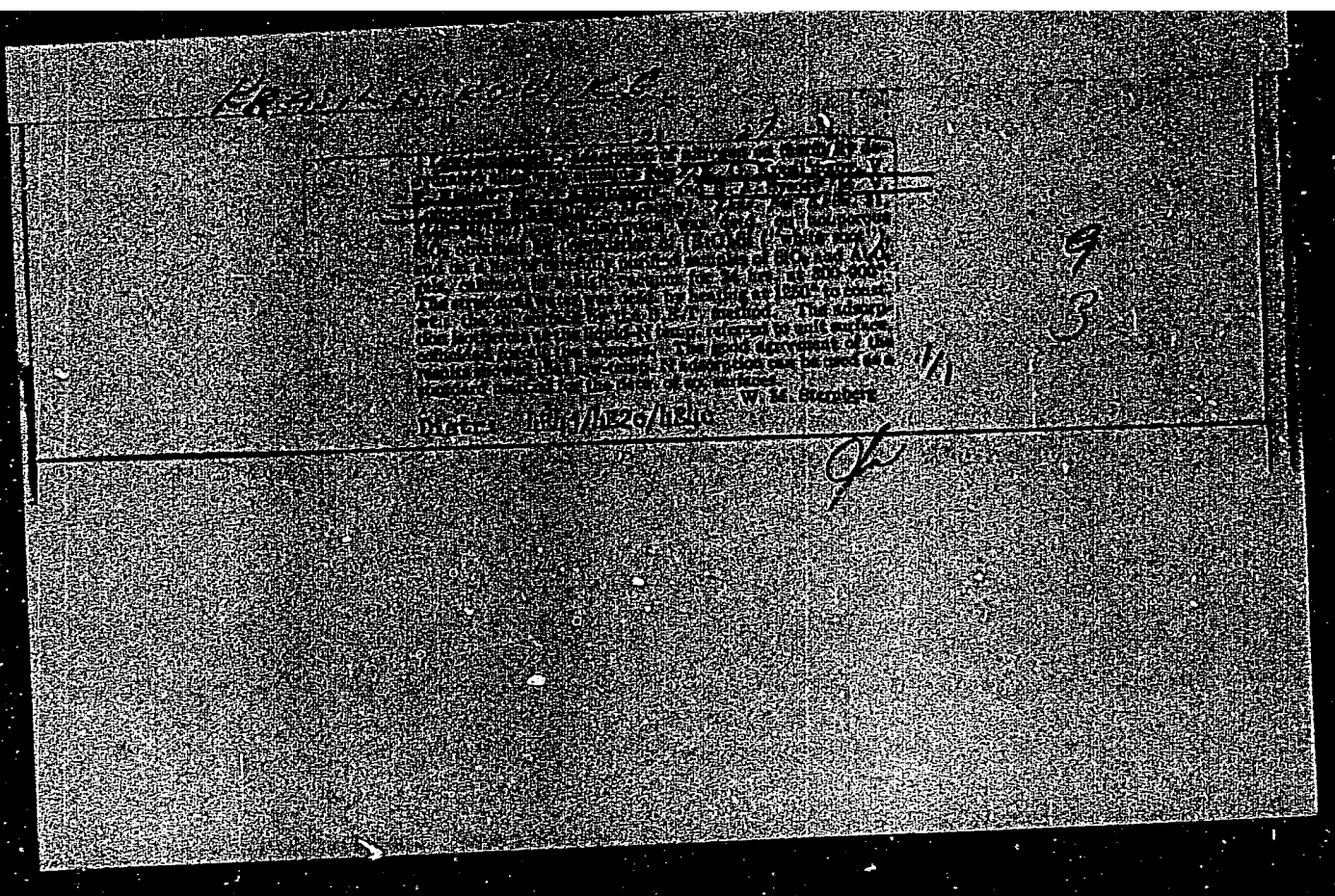
Abs Jour : Referat Zhur - Khimiya, No 4, 1957, 11393

ption of heterogeneity of SG surface. Q_3 is almost not dependent on H_2O . The conclusion is arrived at that most of the earlier data on heat of wetting of SG are not mutually comparable since no account was taken of the correlation between Q and θ H_2O and the nature of porosity of SG (see RZhKhim, 1956, 77773)

IL'IN, B.V.; KISELEV, V.F.; KRASIL'NIKOV, K.G.

Effect of the surface nature of silica on its absorption
properties. Part 1. Vest. Mosk. un. Ser. mat., mekh., astron.
fiz., khim. 12 no. 6:35-50 '57. (MIRA 11:10)

1. Kafedra obshchey fiziki dlia khimicheskogo fakul'teta Moskovskogo
gosudarstvennogo universiteta.
(Silica)
(Absorption)



20-114-3-35/60

AUTHORS: Yegorov, M. M., Yegorova, T. S., Kiselev, V. F.,
Krasil'nikov, K. G.

TITLE: The Adsorption of Water Vapors on Silica Gels Hydrated to
Varied Degrees (Adsorbtsiya parov vody na silikagelyakh razlichnoy
stepeni gidratatsii)

PERIODICAL: Doklady Akademii Nauk SSSR, 1957, Vol. 114, Nr 3, pp 579-582 (USSR)

ABSTRACT: As is known, the adsorption of water vapors on silica gels is
characteristic by some specific properties. Some previously
published scientific papers have investigated in detail the
irreversible adsorption of water vapors which is connected
with an additional hydration of the silica-gel surface in the
process of adsorption. Other investigations reached the con-
clusion that the isotherm of the adsorption of water vapors,
depending on the degree of the dehydration of the silica-gel
surface and of porous glasses, is transformed from a convex
into a concave line, the latter corresponding to a hydro-
phobic surface. There exist different divergences in computing
the specific surfaces of silica gels from the isothermal lines.
None of the authors of the above-mentioned scientific papers

Card 1/4

20-114-3-35/60

The Adsorption of Water Vapors on Silica Gels Hydrated to Varied Degrees

conducted the chemical analysis of the surfaces of the silica gels and of porous glasses. This task was now performed by the authors of the paper under review. Figure Nr 1 of the paper under review represents the isotherms of the water vapors on the initial silica gels and also the curves of distribution - as computed from the desorption branches - of the pore volume with respect to their effective diameter taking into account the thickness of the adsorbed film. Figure Nr 2 contains the initial segments of the primary vapor adsorption on all samples of silica gels, computed for 1 m^2 of the surface. It can be seen from figure Nr 2A that the isotherms of the three initial samples, worked at 300 degrees centigrade, are placed in such a way that p/p_s being the same, the adsorption decreases with a decrease in the degree of hydration of the surface, and this corresponding to the observed reduction in heat of the water moistening of the same samples. The state attained at the water adsorption at the thermally dehydrated surfaces are not equilibrated, as far as in this case the process of hydration of the surface can take place. However, in the monomolecular range at small p/p_s this process is very slow. Therefore it is possible to consider the isotherms of the figure Nr 2A of the silica gel samples K-2, annealed at high tempera-

Card 2/4

20-114-3-35/60

The Adsorption of Water Vapors on Silica Gels Hydrated to Varied Degrees

tures, as equivalent from the point of view of adsorption. For this purpose, however, one has to neglect the slight modification of the surface hydration during the process of establishing the adsorption equilibrium. If these isotherms are compared with the previous ones, it can be seen that, depending on the surface hydration, they change their form and become concave. It is furthermore observed that in this context the capacity of adsorption of the silica gel decreases. Quite a number of assumptions - as found in relevant scientific literature - on the mechanism of adsorption of water vapors on silica gel and on the hydration of its surface, are in contradiction to each other; these assumptions are based on adsorption data and also on the investigation of the infrared spectra of the surface layer. In order to clarify these questions, additional research is necessary, namely study of adsorption linked with spectroscopic investigations. There are 2 figures, 1 table, and 20 references, 14 of which are Slavic.

Card 3/4

20-114-3-35/60

The Adsorption of Water Vapors on Silica Gels Hydrated to Varied Degrees

ASSOCIATION: Moscow State University imeni M. V. Lomonosov
(Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova)

PRESENTED: December 14, 1956, by M. M. Dubinin, Member of the Academy

SUBMITTED: December 10, 1956

Card 4/4

AUTHORS: Bonetskaya, A. K., Krasil'nikov, K. G. 20-114-6-33/54

TITLE: The Adsorption of Aliphatic Alcohols From Solutions on Silica Gel and White Soot
(Adsorbtsiya alifaticheskikh spirtov iz rastvorov na silikagele i beloy sazhe).

PERIODICAL: Doklady AN SSSR, 1957, Vol. 114, Nr 6, pp. 1257-1260 (USSR)

ABSTRACT: From earlier papers (references 1,2,11) follows that the adsorbed maximum amount of alcohols and acids of the homologous series with an increase in the hydrocarbon radical decreases to porous hydrophile adsorbents. In this connection the highest adsorption value in the case of comparatively coarse-grained adsorbents remains constant (references 3,11). The reduction of absorption in the homologous series was explained by a volume-interaction in solutions (different solubility of the terms of the homologous series, reference 1) or ascribed to the influence of the porous structure of the adsorbent (references 3,5). In the case of the nonporous hydrophile adsorbents it could be expected (references 2-4) that the adsorbed maximum quantity of the alcohols and acids of the homologous series was constant. The adsorption measurements

Card 1/5

The Adsorption of Aliphatic Alcohols From Solutions on
Silica Gel and White Soot

20-114-6-33/54

on nonporous hydrophobe adsorbents for unlimitedly soluble alcohols and acids from aqueous solutions (references 6,7) show that the limit of adsorption is shifted upward with the number of carbon atoms in the molecule. Comparisons of adsorption isothermal lines for hydrophile adsorbents are, as far as is known, absent. In the present work the authors performed the adsorption of a number of normal aliphatic alcohols from solutions in CCl_4 on 2 silica samples of different structure. Silica gel KSK-2 (reference 9) was the porous sample. So-called white soot was used as nonporous sample. The samples were sharply different in their structure, but possessed a practically equal hydrated surface. As follows from figure 1 A, the results obtained from both samples are qualitatively not different from each other, for in both cases the maximum value of adsorption decreases with the lengthening of the carbon chain. This value is shifted into the domain of smaller equilibrium concentrations (reference 12). If it is assumed that this decrease can be effected by the competition of the solvent, the adsorption of the latter should increase with

Card 2/5

The Adsorption of Aliphatic Alcohols From Solutions on
Silica Gel and White Soot

20-114-6-33/54

a decrease of the adsorption of alcohol, i.e. on transition to the higher alcohols. The authors tried an extrapolation of the inclined part of the adsorption isotherm of octyl-alcohol on silica gel to the domain of high concentrations, so that a value was obtained which lies close to the concentration of pure alcohol. Due to the inaccuracy of such an extrapolation additional tests on the adsorption of CCl_4 and its solutions in octylalcohol were carried out. But no perceptible modifications of adsorption were determined in the domain of concentrations in question here. From the isothermal lines obtained the authors calculated the values of the total content of the adsorbed substance in the surface layer α (table 1). Thus the decrease in adsorption of the alcohols with a lengthening of the hydrocarbon chain cannot be ascribed to the competition of the solvent. The authors try to explain this phenomenon by a special mode of orientation of the alcohol-molecules in the solid surface layer so that the chains predominantly "lie" on the surface. In such a case the surface taken up by the molecule will be dependent on the above-mentioned length of chain. For determining the

Card 3/5

The Adsorption of Aliphatic Alcohols From Solutions on
Silica Gel and White Soot

20-114-6-33/54

influence of the porous structure upon the maximum value of absorption the obtained results were related to 1 m. of the surface. The adsorption isothermal lines of methyl alcohol on both adsorbents agree (figure 1 B). The adsorption on the porous sample increases with increasing number of carbon atoms. According to table 1 and figure 2 the maximum volume of the adsorption of all investigated alcohols changes little on nonporous soot. There are 2 figures, 1 tables, and 16 references, 11 of which are Slavic.

ASSOCIATION: Moscow State University imeni M. V. Lomonosov,
(Moskovskiy gosudarstvennyy universitet im. M. V.
Lomonosova)

Card 4/5

The Adsorption of Aliphatic Alcohols From Solutions on
Silica Gel and White Soot

20-114-6-33/54

PREPARED BY: N. M. Dubinin, Academician
APPROVED FOR RELEASE: Monday, July 31, 2000

CIA-RDP86-00513R0008

100

SUBMITTED: December 29, 1956

Card 5/5

KRASIL'NIKOV, K.G.

AUTHORS:

Krasil'nikov, K. G., Kiselev, V. F., Sysoyev, Ye. A. 20-6-27/42

TITLE:

Nature of the Surface of a Dehydrated Silicagel
(K voprosu o prirode poverkhnosti degidratirovannogo
silikagelya)

PERIODICAL:

Doklady AN SSSR, 1957, Vol. 116, Nr 6, pp. 990-993 (USSR)

ABSTRACT:

The authors carried out quantitative measurements of the adsorption of nitrogen and oxygen on silicagels which were dehydrated in high vacuum. The adsorption was measured by means of the volum method. The silicagel test piece was introduced into a quartz ampule and annealed after previous draining at 300°C at an assumed temperature. Then the prepared portion of the gas to be investigated was introduced into the ampule and the corresponding measurements were carried out at 20° C. Nitrogen is not adsorbed under these conditions within the accuracy of measurement. With oxygen, the surface of silicagel dehydrated in vacuum at temperatures of 300 to 900° C adsorbs the oxygen to a considerable extent. Hereby the quantity of adsorbed oxygen grows with an increase of the annealing temperature. The effect of a short-wave radiation and the thermic dehydration in the final effect apparently lead to the same properties of the surface. The authors further investigated the

Card 1/2

~~Notes of the~~ Surface of a Dehydrated Silicagel

20-6-27/42

heats of wetting of the silicagels with water in which case these silicagels were previously annealed in vacuum at various temperatures up to 800° C. The data obtained during this operation are summarized in a table. The two silicagels investigated here, produce after annealing in vacuum a greater heat of wetting than the same test pieces annealed in air. On the surface of the silicagel dehydrated in vacuum, centers with higher activity of adsorption than with the OH-groups are formed. The results obtained in this case agree with the measurements of other authors (reference 11,12). There are 2 figures and 12 references, 8 of which are Slavic.

ASSOCIATION: Moscow State University im. M. V. Lomonosov
(Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova).

PRESENTED: May 4, 1957, by M. M. Dubinin, Academician.

SUBMITTED: May 26, 1957

AVAILABLE: Library of Congress

Card 2/2

AUTHORS: Yegorov, M.M., Yegorova, T.S., Kiselev, V.F., SOV/55-58-1-27/33
and Krasil'nikov, K.G.

TITLE: Influence of the Nature of the Silica Gel Surface on the
Adsorption of the Methyl Alcohol Vapors (Vliyaniye prirody poverkh-
nosti silikagelya na adsorbtsiyu parov metilovogo spirta)

PERIODICAL: Vestnik Moskovskogo universiteta, Seriya fiziko-matematicheskikh i
yestestvennykh nauk, 1958, Nr 1, pp 203-207 (USSR)

ABSTRACT: The paper is written under the leading of Professor B.V. Il'in
and contains the results of a detailed measuring of methyl
alcohol vapors which in the monomolecular range have been adsorbed
at the surface of the silica gel. Before the experiment, the
surface of the silica gel was submitted to the influence of
saturated water vapor up to 48 hours. The results are collected
in a table and two figures.
There are 15 Soviet references.

ASSOCIATION: Kafedra obshchey fiziki dlya khimicheskogo fakul'teta (Chair of
General Physics of the Department of Chemistry)

SUBMITTED: May 3, 1957

Card 1/1

5(4)

AUTHORS:

Il'in, B.V., Kiselev, V.F., and
Krasil'nikov, K.G.

SOV/55-58-2-31/35

TITLE:

Heat of Wetting of the Silica Gels of Different Degrees of
Hydration (Teploty smachivaniya silikageley razlichnoy
stepeni gidratatsii)

PERIODICAL:

Vestnik Moskovskogo Universiteta. Seriya matematiki, mekhaniki,
astronomii, fiziki, khimii, 1958, Nr 2, pp 223-232 (USSR)

ABSTRACT:

The paper contains the results of a systematic investigation
of the heat of wetting of different kinds of silica gels.
The wetting of the surface was carried out by water, n-propyl-
alcohols and n-heptane. The structural water content of the
silica gel was taken into account. Already known properties
were essentially confirmed. The opinion of A.V. Kiselev and
his collaborators [Ref 9-16] was not confirmed according to
which the unit of the surface of the silica gel possesses
certain "absolute" energetic properties. This is not the
case: The properties of the surface essentially depend on
the preceding treatment (annealing etc), i.e. on the bound
water content of the surface layer.
There are 6 figures, and 25 references, 15 of which are Soviet,

Card 1/2

Heat of Wetting of the Silica Gels of Different
Degrees of Hydration

SOV/55-58-2-31/35

7 American, 1 English, and 2 French.

ASSOCIATION: Kafedra obshchey fiziki dlya khimicheskogo fakul'teta
(Chair of General Physics of the Faculty of Chemistry)

SUBMITTED: April 14, 1957

Card 2/2

AUTHORS: Kiselev, V.F., Krasil'nikov, K. G. SOV/76-32-6-45/46
TITLE: The Specific Character of the Adsorption of Phenol by Silicagel
From Heptane Solutions (Osobennosti adsorbtsii fenola iz
rastvorov v geptane silikagelem)
PERIODICAL: Zhurnal fizicheskoy khimii, 1958, Vol. 32, Nr 6, pp 1435-1436
(USSR)
ABSTRACT: In a previous paper it was found that the initial domain of
the adsorption isothermal line shows a steplike character;
more accurate measurements in this field showed that great
changes of the integral heat adsorption according to the con-
centration take place. In connection with observations made by
other authors it turned out to be interesting to carry out
parallel experiments of the adsorption of phenol from solutions
for purposes of investigating the adsorption isothermal line
on the one hand and the heat of wetting of the same solutions
on the same silicagel on the other hand. The authors used a
coarse-pored silicagel KSK-1, the methods of measurement re-
maining the same as in the previous paper. The experimental
results obtained do not yet permit the interpretation of the
observations made, however, the authors put forward some ex-

Card 1/3

The Specific Character of the Adsorption of Phenol by SOV/76-32-6-45/46
Silicagel From Heptane Solutions

planations from which it may be seen that the phenomena are due to the complicated process of the filling of the surface of the adsorbent, which according to its properties is inhomogeneous, with the molecules of the substance to be adsorbed. It was found that the change of the chemical nature of the surface of the adsorbent caused by different ways of treatment (e.g. dehydration) can lead to the occurrence of steps in the isothermal line or to their removal, respectively. In order to be able to explain the occurrence of maxima and minima found on the isothermal line of the heat of wetting, or to find a possible connection with the step-phenomenon on the adsorption isothermal line more experiments will have to be carried out. There are 2 figures and 8 references, 7 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M.V.Lomonosova
(Moscow State University imeni M.V. Lomonosov)
SUBMITTED: December 11, 1957

Card 2/3

The Specific Character of the Adsorption of Phenol by Silicagel From Heptane Solutions SOV/76-32-6-45/46

1. Phenols--Adsorption 2. Colloidals--Adsorptive properties

Card 3/3

5(4)

AUTHORS:

Yegorov, M. M., Krasil'nikov, K. G.,
Kiselev, V. F.

SSV/76-32-10-53/39

TITLE:

The Influence of the Nature of Silica Gel and Quartz
Surfaces on Adsorption Properties (Vliyaniye prirody poverkh-
nosti silikagelya i kvartsa na ikh adsorbtsionnyye svoystva)
I. Investigations of the Hydration of the Silicon Dioxide
Surface (I. Issledovaniya gidratatsii poverkhnosti
kremnezema)

PERIODICAL:

Zhurnal fizicheskoy khimii, 1958, Vol 32, Nr 10,
pp 2448 - 2454 (USSR)

ABSTRACT:

Of late the presence of hydroxyl groups on silicon
dioxide surfaces was found in investigations (Refs 8-13).
The present paper deals in detail with investigations of
the degree of hydration in dependence on the annealing
in 7 different SiO_2 samples. The silica gel K&K was
carefully purified; silica gel K-2 was obtained by a
distillation of SiCl_4 according to a method mentioned
(Ref 3), and after storing under water it was termed
silica gel K-3. "White root" and ground quartz (sample
BS-1) were used as non-porous samples. The determinations

Card 1/4

The Influence of the Nature of Silica Gel and Quartz Surfaces on Adsorption Properties. I. Investigations of the Hydration of the Silicon Dioxide Surface SSN/76-32-10-33/39

of the specific surfaces of the samples were carried out according to the BET method by means of nitrogen vapors. All silica gel samples used belong to the type of coarsely porous adsorbents (Ref 16). Diagrams of the function of the water content versus the annealing temperature of the silica gels KSK-1, KSK-2, K-2 and K-3 are given using data by Shapiro and Weiss (Veys) (Ref 14) as well as by Bastick (Bastik) (Refs 4, 17). The standard temperature for treating the samples was chosen to be 300°. The results show that the content of the water of constitution as related to the surface unit is different for various silica gels. In the case where the samples were treated exactly the same but a different specific surface was present no surfaces with the same degree of hydration could be obtained, which proves the incorrectness of the data mentioned in reference 21. On storing the samples in water it was found that the amount of water of constitution on the surface increased sharply. However, those samples treated

Card 2/4

The Influence of the Nature of Silica Gel and Quartz Surfaces on Adsorption Properties. I: Investigations of the Hydration of the Silicon Dioxide Surface SOV/76-32-10-33/39

at 300° after storing always had smaller amounts than those without any temperature treatment. The process of dehydration and secondary hydration was irreversible under the conditions present. The degree of hydration depends on the crystal chemical properties of the surface structure and is determined by the valence number of the surface atoms that are loosely bound. The surface hydration of all samples investigated treated under the same conditions was different and amounted to a maximum of 4,78 μ M/m². A paper by Eiler (Iler) (Ref 4) is mentioned; the authors thank B.V.Ill'in. There are 3 figures, 1 table, and 23 references, 12 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im.M.V.Lomonosova
(Moscow State University imeni M.V.Lomonosov)

Card 3/4

The Influence of the Nature of Silica Gel and Quartz Surfaces on Adsorption Properties. I. Investigations of the Hydration of the Silicon Dioxide Surface SO7/76-32-10-33/39

SUBMITTED: June 5, 1957

Card 4/4

SCV/76-32-11-25/32

5(4)

AUTHORS:

Yegorov, M. M., Yegorova, T. S., Krasilnikov, K. G.,
Kiselev, V. F.

TITLE:

The Effect of the Nature of the Silica Gel and Quartz Surface
on Its Adsorption Properties (Vliyaniye prirody poverkhnosti
silikagelya i kvartsa na ikh adsorbtsionnyye svoystva) II.
Adsorption of Steam, Methyl Alcohol and Nitrogen on Silica
Gel of Different Degrees of Hydration (II. Adsorbtsiya parov
vody, metilovogo spirta i azota na silikagelyakh razlichnoy
stepeni gidratatsii)

PERIODICAL:

Zhurnal fizicheskoy khimii, 1958, Vol 32, Nr 11, pp 2624-2633
(USSR)

ABSTRACT:

Silica gel samples and non-porous "white soot" described in
the previous paper were used. The measurements of the adsorp-
tion were carried out according to the gravimetric method. It
was found (Fig 1) that with samples treated at 300°C the ad-
sorption (at constant p/p_s) decreases with a decrease of the
degree of hydration of the surface. The different adsorbability
of the investigated silica gels is not due to their structure

Card 1/3

SOV/76-32-11-25/32

The Effect of the Nature of the Silica Gel and Quartz Surface on Its Adsorption Properties. II. Adsorption of Steam, Methyl Alcohol and Nitrogen on Silica Gel of Different Degrees of Hydration

but to the chemical nature of the surface (their degree of hydration). It is assumed that the hydroxyl groups with water molecules can form hydrogen compounds on the surface (Ref 12), and thus act as adsorption centers. Contradicting data given by other authors on the adsorption centers mentioned above (Refs 15,16) are explained by a different technique of investigation. As the hydration of the surface of the investigated samples is different the adsorption properties of the surface with respect to the molecules capable of forming hydrogen compounds with hydroxyl groups are also different. Measurements carried out of the surface of hydrated KSK-1 samples occupied by water molecules showed that within the range of p/p_s from 0.1 to 0.3 the value ω changes from 39 to 22.5 \AA^2 and thus is considerably higher than that given in publications (10.6 and 14.8 \AA^2) (Refs 20-22). As the adsorption properties are functions of several factors (crystallography of the sample, chemical composition etc.) they cannot be called "absolute" properties ("absolute" isothermal lines). The authors thank M. M. Dubinin and B. V. Il'in.

Card 2/3

SOV/76-32-11-25/32

The Effect of the Nature of the Silica Gel and Quartz Surface on Its
Adsorption Properties. II. Adsorption of Steam, Methyl Alcohol and Nitrogen
on Silica Gel of Different Degrees of Hydration

There are 8 figures and 29 references, 22 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University imeni M. V. Lomonosov)

SUBMITTED: June 5, 1957

Card 3/3

AUTHORS: Yegorov, M. M., Zarif'yants, Ye. A., 027/86-100-2-26/63
Kiselev, V. F., Krasil'nikov, K. G.

TITLE: The Adsorption Properties of Alumina-Silicate Catalysts and Their
Dependence Upon Composition (Adsorbtsionnyye svoystva aluminosili-
katnykh katalizatorov i ikh zavisimost' ot sostava)

PERIODICAL: Doklady Akademii nauk SSSR, 1958, Vol. 120, Nr 2,
pp. 326 - 329 (USSR)

ABSTRACT: In some previous papers (Refs 1-4) it was shown that the adsorption
properties per unit of surface with respect to water and ethyl
alcohol molecules are to a considerable extent dependent upon the
degree of hydration of the surface. It would be of interest to
extend such investigations to a number of alumina-silicates of
varying composition. In the first stage of these studies the authors
investigated the adsorption of steam and of methyl alcohol vapors
and the heat necessary to wet the synthetic alumina-silicate compounds.
The catalysts had a content of 15% (Gudri catalyst), of 30% and of
50% of Al_2O_3 . The measurements of adsorption were carried out in

Card 1/3 a calorimeter with constant heat exchange. A diagram gives the

The Adsorption Properties of Alumino-Silicate Catalysts and Their Dependence Upon Composition

SOV/20-121-8-78/63

Function of the heat required for wetting by water versus the content of crystal water for all alumino-silicate under investigation. These curves exhibit maxima which reproduce the thermal pre-treatment of the samples at 200-300°. The comparatively high content of crystal water is of interest, in particular in the samples with a high Al_2O_3 content. The heats of wetting differ by about the double between silicagel and alumino-silicate with a low Al_2O_3 content (15%) even with a similar hydration of the surface.

The same samples were also used for the determination of the isothermal lines of the adsorption of steam and of methyl alcohol vapors. The desorption isothermal lines of all samples are considerably below the adsorption isothermal lines. Silicagel, however, did not show such a behaviour. The structure of alumino-silicates is similar to that of silica, its surface, however, is more inhomogeneous. Investigation of the adsorption mechanism cannot be limited to the local adsorbed molecules with active centers, and their topography and their concentration must be taken into account. In conclusion the authors express their gratitude

Card 2/3

The Adsorption Properties of Alumo-Silicate
Catalysts and Their Dependence Upon Composition

SOV/2c-12c-2-28/63

to B.V.Ilin for the help rendered by him in this work and to
K.V.Topchiyev for his unabated interest and for his furnishing the
samples. There are 3 figures, 1 table, and 16 references, 15 of
which are Soviet.

ASSOCIATION: Fizicheskiy fakul'tet Moskovskogo gosudarstvennogo universiteta
im.M.V.Lomonosova (Dept. of Physics of the Moscow State University
imeni M.V.Lomonosov)

PRESENTED: January 15, 1958, by M.M.Dubinin, Member, Academy of Sciences,
USSR

SUBMITTED: January 6, 1958

1. Aluminum silicate catalysts--Adsorptive properties
2. Aluminum silicate catalysts--Physical properties

Card 3/3

5(4)

AUTHORS:

Yegorova, T. S., Kiselev, V. F.,
Krasil'nikov, K. G.

SOV/20-123-6-28/50

TITLE:

The Differential Heats of the Adsorption of Water Vapors on
Silica Gels of Different Hydration (Differentsial'nyye toploty
adsorbtsii parov vody na silikagelyakh razlichnoy gidratatsii)

PERIODICAL:

Doklady Akademii nauk SSSR, 1958, Vol 123, Nr 6, pp 1060-1063
(USSR)

ABSTRACT:

No reliable data have hitherto been published on the dependence of the differential adsorption heats of water vapors on the filling up of the surface. In the present paper the silica gels K - 2 and KSK - 3 were used. The characteristic data of the adsorption on these samples are given in a table. The adsorption heats of the vapors were measured in a calorimeter similar to that described by reference 7; the wetting heats were measured in a calorimeter with constant heat exchange. The water vapors were adsorbed at constant vapor pressure. The authors investigated the initial domains of isothermal lines and of the differential adsorption heats of water vapors in various silica gels by means of two methods. A diagram shows the wetting heats

Card 1/3

The Differential Heats of the Adsorption of
Water Vapors on Silica Gels of Different Hydration

SOV/20-123-6-28/50

as functions of the previously adsorbed quantity of water. In a previous paper (Ref 1) homogeneous large-pore adsorbents were investigated within the domain of adsorption up to the beginning of capillary condensation. The results obtained by calculating the differential adsorption heat as a function of specific adsorption are shown in form of a diagram. The adsorption heats for the silica gel K - 2 - 300°, which were determined by means of direct calorimetical measurements, agree well with the theoretically calculated curves. The initial values of water adsorption on silica gel KSK are within the interval of 15 - 20 kcal/mol. At low degrees of filling the adsorbed molecules form 3 or even 4 hydrogen bonds with the hydroxyls of the surface. Part of the molecules is probably adsorbed within this domain on centers of higher energy. In the case of one and the same degree of filling the differential heats decrease with a decreasing degree of hydration of the surface. Also the differential entropy of water vapor adsorption

Card 2/3

The Differential Heats of the Adsorption of
Water Vapors on Silica Gels of Different Hydration

SOV/20-123-6-28/50

decreases with increasing surface hydration of the silica gels.
The authors thank B. V. Il'in for his interest in this work
and for discussing results. There are 3 figures, 1 table, and
14 references, 10 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University imeni M. V. Lomonosov)

PRESENTED: May 28, 1958, by M. M. Dubinin, Academician

SUBMITTED: May 26, 1958

Card 3/3

5(3), 5(4)
AUTHORS:

SCV/156-59-1-12/54

Zarif'yants, Yu. A., Kapitonova, N. V., Kiselev, V. F.,
Krasil'nikov, K. G.

TITLE:

The Adsorption of Benzene Vapors on Aluminosilicates of Various Composition (Adsorbtsiya parov benzola na al'yumossilikatakh razlichnogo sostava)

PERIODICAL:

Nauchnyye doklady vysshey shkoly. Khimiya i khimicheskaya tekhnologiya, 1959, Nr 1, pp 48 - 51 (USSR)

ABSTRACT:

The insertion of AlO_4 tetrahedrons in the structure of silica leads to a variation of the hydrated as well as unhydrated sectors of the surface. Thus also the adsorption properties vary during the transition from pure silica to aluminosilicates of various composition. Aluminosilicates with a content of 15% and 30% Al_2O_3 as well as the aluminogel AT and silica gel K-2 were investigated. The isothermal lines of adsorption are given in diagrams. The initial sections (in enlarged reproduction) lie

Card 1/3

The Adsorption of Benzene Vapors on Aluminosilicates
of Various Composition

SOV/156-59-1-12/54

on a curve, and the adsorption rises with increasing Al_2O_3 content. This cannot be explained by an increase of the adsorption potential in the pores. The adsorption of aluminogel is higher than that of equally porous aluminosilicate with 15% Al_2O_3 and of more fine-porous silica gel. The variation of the adsorptive capacity seems to depend on changes of the surface structure. This will be investigated with nonporous adsorbents in a future work. V. T. Bykov (Ref 8) assumed that the so-called "absolute" adsorption properties of the surface of silica and aluminosilicates are equal and extended this statement to various kinds of adsorbents. This is a false presumption, based on unfounded presuppositions. Actually, a function must be effective here which depends just on the specific properties of the surface of the individual adsorbents. The range, for instance, which is occupied by a benzene molecule on silica gel is larger than that on the aluminogel. Gratitude is expressed to B. V. Il'in for his assistance in this work. There are 2 figures and 16 refer-

Card 2/3

The Adsorption of Benzene Vapors on Aluminosilicates
of Various Composition

SOV/156-59-1-12/54

ences, 14 of which are Soviet.

ASSOCIATION: Kafedra obshchey fiziki Moskovskogo gosudarstvennogo univer-
siteta im. M. V. Lomonosova (Chair of General Physics of
Moscow State University imeni M. V. Lomonosov)

SUBMITTED: July 10, 1958

Card 3/3

SOV/153-2-3-9/29

5(4)

AUTHORS:

Yegorov, M. M., Kiselev, V. F., Krasil'nikov, K. G.,
Simanov, Yu. P.

TITLE:

The Influence of the Phase Composition of the Adsorbents
in the System $Al_2O_3 \cdot H_2O$ on Their Surface Properties

PERIODICAL:

Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya
tekhnologiya, 1959, Vol 2, Nr 3, pp 360-365 (USSR)

ABSTRACT:

Cherenkov aluminum oxide from the laboratory of K. V. Topchiyeva
khimicheskii fakul'tet MGU (Chemical Department of Moscow State
University) was used for the investigation. The dehydration at
different temperatures was investigated (Fig 1). Phase investi-
gations were carried out by X-ray methods with cameras of the
type RDK-57 and with X-ray tubes of the type BSV. The samples
were tempered at different temperatures and the wetting heat
was determined (Table). The results are - referred to 1 g
oxide - represented in diagrams (Fig 2). A second representation
is given with respect to the surface unit (Fig 3). A dependence
between the structural water and the wetting heat per surface
unit was found (Fig 4). The phase change and the change of the
degree of wetting of the surface causes a sharp change of the

Card 1/2

The Influence of the Phase Composition of the SOV/153-2-3-9/29
Adsorbents in the System $Al_2O_3 - H_2O$ on Their Surface Properties

surface properties. The authors thank K. V. Topchiyeva and
B. V. Il'in for their assistance in the investigations.
There are 4 figures, 1 table, and 10 references, 7 of which
are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet imeni M. V. Lomonosova -
Kafedra fiziki (Moscow State University imeni M. V. Lomonosov -
Chair of Physics)

SUBMITTED: April 24, 1958

Card 2/2

5(4)

SOV/76-33-1-11/45

AUTHORS: Yegorov, M. M., Kiselev, V. F., Krasil'nikov, K. G., Murina, V. V.

TITLE: The Effect of the Surface Nature of Silica Gel and Quartz on Their Adsorption Properties (Vliyaniye prirody poverkhnosti silikagelya i kvartsa na ikh adsorbtsionnyye svoystva) III. Heats of Wetting of Silicon Dioxide With Various Liquids (III. Teploty smachivaniya kremnezema razlichnymi zhidkostyami)

PERIODICAL: Zhurnal fizicheskoy khimii, 1959, Vol 33, Nr 1, pp 65-73 (USSR)

ABSTRACT: In connection with previous papers the effect of the hydration of the surface of silicon dioxide on the adsorption energy of water and methanol in the form of heat of wetting (HW) is investigated. HW was determined in several SiO_2 samples with water, methanol, n-propanol, and n-heptane in dependence on the hydration degree of the surface. Data on the HW of the silica gels KSK with water were taken from M. M. Yegorov's thesis (Ref 18). The HW was measured by means of a calorimeter with a temperature sensitivity of $5 \cdot 10^{-5}^\circ\text{C}$. A table of the investigated silica gels with the HW obtained for water is given. An investigation of the effect of the glowing temperature on the HW

Card 1/3

SOV/76-33-1-11/45

The Effect of the Surface Nature of Silica Gel and Quartz on Their Adsorption Properties. III. Heats of Wetting of Silicon Dioxide With Various Liquids

(Fig 1) showed that a glowing temperature of 200-300°C the function curves pass through a maximum. An increase in the glowing temperature up to 1000°C resulted in a surface decrease, e. g. in silica gel K-2, of several m^2/g . A treatment at 300°C is considered the standard. Here, the dependence of the HW on the hydration of the surface is expressed by a straight line. A wetting of thermally dehydrated samples with water results in the formation of hydration heat. A hydrated quartz surface differs qualitatively from a corresponding silica gel surface which can be explained by the closer packing of the hydroxyl groups (in quartz); however, investigations have still to be carried out in this respect (e. g. according to the method of the core-paramagnetic resonance). The HW of methanol does not depend on the porosity of the silica gels, which is the case with n-propanol and n-heptane. In the case of partly dehydrated surfaces a greater HW is obtained by the use of methanol than by that of water which can be explained by the effect of the methyl group in the adsorption. The results of the investigations show that the HW

Card 2/3

SOV/76-33-1-11/45
The Effect of the Surface Nature of Silica Gel and Quartz on Their Adsorption Properties. III. Heats of Wetting of Silicon Dioxide With Various Liquids

of the silica gel with water and methanol depends essentially on the hydration degree of the surface which is not the case with n-heptane. The authors thank B. V. Il'in and G. I. Aleksandrova. There are 3 figures, 1 table, and 22 references, 14 of which are Soviet.

ASSOCIATION:

Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University imeni M. V. Lomonosov)

Card 3/3

05806

SOV/76-33-10-4/45

5(4)

AUTHORS:

Yegorov, M. M., Kiselev, V. F., Krasil'nikov, K. G.

TITLE:

On the Problem of the Adsorptive Power of a Unit of the Quartz Surface

PERIODICAL:

Zhurnal fizicheskoy khimii, 1959, Vol 33, Nr 10, pp 2141-2144 (USSR)

ABSTRACT:

Since the quantity of adsorbed OH groups depends on the number of free corners of the SiO_4 tetrahedron which project into the surface of the silica-gel skeleton, it was assumed (Refs 2-4) that differences in the degree of hydration of silica gels (Refs 1-4) is connected with the manner in which the tetrahedron is packed (in dependence on the conditions of silica-gel preparation). The adsorptive properties of samples of amorphous silicon dioxide of various origin (silica gels and quartz glass) were therefore compared with those of quartz samples since the latter has the densest packing of SiO_4 tetrahedrons.

The authors investigated powder samples obtained by grinding (carried out by L. A. Feygin), crystalline quartz and transparent quartz glass. The samples were ground in dry state as well as under the addition of water. The adsorptive properties of the samples are listed (Table: quartz, Kv-1, -2, -3 samples, quartz glass, sample KS-1 and the silica gels KSK-1

Card 1/2

05806

SOV/76-33-10-4/45

On the Problem of the Adsorptive Power of a Unit of the Quartz Surface

and K-2). Comparison of the adsorption isothermal lines (Fig) shows that the adsorptive power of quartz depends on the conditions under which the samples were pulverized. The adsorption isothermal lines of the samples which were ground in dry state or with a small addition of water attain considerably higher values than those of samples ground in wet state. The isothermal lines of water vapor adsorption of samples (ground under the same conditions) obtained from crystalline quartz Kv-3 and quartz glass KS-1 (Fig 1) indicate that the isothermal line of isotropic quartz glass attains higher values than that of crystalline quartz. Accordingly, results do not confirm the view of A. V. Kiselev (Refs 11, 12) as to the same adsorptive properties of crystalline and amorphous silicon dioxide. In conclusion, the authors thank K. V. Chmutov. There are 2 figures, 1 table, and 12 references, 9 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University imeni M. V. Lomonosov)

SUBMITTED: February 26, 1958

Card 2/2

SOV/20-125-4-40/74

5(4)
AUTHORS:

Bakayev, V. A., Kiselev, V. F.,
Krasil'nikov, K. G.

TITLE:

The Reduction of the Melting Temperature of Water in the
Capillaries of a Porous Body (Ponizheniye temperatury plavleniya
vody v kapillyarakh poristogo tela)

PERIODICAL:

Doklady Akademii nauk SSSR, 1959, Vol 125, Nr 4, pp 831-834
(USSR)

ABSTRACT:

From the data concerning the phase composition of an adsorbed substance as a function of temperature it is possible to determine the quantitative characteristic of the structure of a porous body by determining not only the radius but also the volume of the capillaries in which the phase transformations take place. The quantity of adsorbed substance in 1 g of the adsorbent melting at the temperature T can be determined from the specific heat of the system adsorbent-adsorbed substance. A more simple, but sensitive method is that of indirect determination of heat capacity by measuring the temperature conductivity λ of the system. The authors carried out these measurements by employing the modified method of "linear temperature increase". The adsorbents used were the silica gels KSK-2,

Card 1/3

The Reduction of the Melting Temperature of Water in the Capillaries of a Porous Body SOV/20-125-4-40/74

KSM-1 and a specimen of a non-porous alumina BS-1. In these samples the isothermal lines of the adsorption of water vapors were measured. Measurements of temperature conductivity were carried out ranging from the temperature of liquid nitrogen to the temperature of 275° K. The dependences of the quantity const/λ on temperature thus determined are shown by a diagram. The theory of capillary condensation shows a connection between the reduction of temperature of the phase transformation and the radius of the capillaries containing the adsorbent substance. A connection between the freezing temperature of water and the radius of the pores can be derived. The points in the diagram $\Delta T = f(10^3/r)$, which were determined for various samples and by various methods, are well suited for a straight line. The method of determining const/λ suggested by the authors makes it possible quickly to determine the substance adsorbed in the porous body. Herefrom it is then possible to determine the curve for the distribution of the volume of the pores over their effective radii. The authors thank L. V. Radushkevich for his interest in this investigation.

Card 2/3

The Reduction of the Melting Temperature of Water in
the Capillaries of a Porous Body

SOV/20-125-4-40/74

There are 3 figures and 8 references, 2 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University imeni M. V. Lomonosov). Institut
fizicheskoy khimii Akademii nauk SSSR (Institute of Physical
Chemistry of the Academy of Sciences, USSR)

PRESENTED: December 24, 1958, by M. M. Dubinin, Academician

SUBMITTED: December 17, 1958

Card 3/3

5 (4), 15 (2)
AUTHORS:

Ganichenko, L. G., Kiselev, V. F.,
Krasil'nikov, K. G.

SOV/20-125-6-29/61

TITLE:

The Influence of the Hydration of the Surface of Silica on the
Adsorption of Aliphatic Alcohols From Solutions (Vliyaniye
gidratatsii poverkhnosti kremnezema na adsorbtsiyu
alifaticeskikh spirtov iz rastvorov)

PERIODICAL:

Doklady Akademii nauk SSSR, 1959, Vol 125, Nr 6,
pp 1277-1280 (USSR)

ABSTRACT:

The influence exercised by the hydration of the surface of
silica is investigated for the adsorption of steam (Ref 1)
and saturated hydrocarbons (Ref 2). In the former case this
influence is considerable, in the latter it is insignificant.
It was therefore of interest to investigate this influence in
the adsorption of alcohols which have both hydroxyl groups
and carbon chains. Measurements were carried out of the
adsorption of methanol-, n-propanol-, n-hexanol, and n-octanol
from carbon tetrachloride solutions. Two samples of non-
porous silica - "white carbon black" - BS-1 and BS-2 were used.
The samples were annealed before the experiments at 300°, one
of the BS-2 samples also at 700°. The results obtained are

Card 1/3

The Influence of the Hydration of the Surface of Silica on the Adsorption of Aliphatic Alcohols From Solutions SOV/20-125-6-29/61

shown by table 1. Figure 1 shows the isothermal lines of adsorption, figure 2 shows the dependence a) of the adsorption maximum, b) of the surface occupied by the adsorbed molecules, c) of the thickness of the adsorption layer, d) of the ratio between the adsorbed molecules and the number of hydroxyl groups on the degree of surface hydration. Whereas methanol is still considerably influenced by the degree of hydration, this influence decreases with an increase of the carbon chain. The adsorption of octanol is not influenced at all. With an increasing length of the carbon chain the behavior of the alcohols thus approaches that of the hydrocarbons. Further, the marked increase in thickness of the adsorption layer of methanol is discussed. It is explained by variation of molecule orientation, which may be caused by a polymorphic transformation due to the thermal treatment of the silica, and leads to steps or discontinuities in the adsorption isothermal line. There are 3 figures, 1 table, and 15 references, 13 of which are Soviet.

Card 2/3

The Influence of the Hydration of the Surface of SOV/20-125-6-29/61
Silica on the Adsorption of Aliphatic Alcohols From Solutions

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
 (Moscow State University imeni M. V. Lomonosov)

PRESENTED: December 30, 1958, by M. M. Dubinin, Academician

SUBMITTED: December 24, 1958

Card 3/3

KRASIL'NIKOV, K. G.

3/042/00/000/009/002/001
3033/3064

AUTHORS: Kamchenko, L. G., Dubinin, M. M., Zvereva, Ye. D.,
Kiselev, V. P., and Krasil'nikov, K. G.

TITLE: Study of the Vapor Adsorption on Microscopic With
Heterogeneous Surface. Communication I. Experiments With
Organically Substituted Silica Gel

PERIODICAL: Izvestiya Akademii nauk SSSR. Otdeleniye Khimicheskikh
nauk, 1960, No. 9, pp. 1535-1545

TEXT: The adsorption of various vapors on methylated coarse-porous silica
gel and a desubstituted sample obtained therefrom is discussed here. The
conditions of investigation were chosen in such a way that an essential
change of the specific surface seemed to be unlikely. Consequently,
mercurial silica gel KGA (MSK) was taken as initial adsorbent. The
adsorbent was purified from iron and other impurities. To study the
gel was repeatedly treated with dichloro acid. The silica gel washed
Then, the vapors were sucked off in vacuum at 10°C, and silica gel washed
with water until the reaction for the chlorine ion was negative. The

Card 1/3

silica gel thus obtained is called C-1 (S-1). A portion of silica gel S-1
was treated with nitric acid vapors at 200°C for 6 h. Thus, the organic
part of the surface was oxidized, the CH₃ radicals substituted by OH
groups, and the substituted silica gel with hydrophobic properties became
hydrophilic. This specimen was called C-2 (S-2). The composition of the
surface of silica gel S-1 and S-2 was determined by an organic analysis.
The analysis was carried out at the same time as the determination of the
weight losses in calcination at 1250°. Table 1 shows the analytical to an
results. It may be seen that the desubstitution of the surface leads to an
increase of its degree of hydration. The specific surface of the specimen
changed by 2%. Nitrogen, cyclohexane, benzene, and water were used as
adsorbates. Figs. 1-4 show the sorption isotherms. In all cases, the iso-
thermal lines of the vapors of these substances are shifted to the right.
Thermal lines for S-1 are shown in these for S-2. Table 2 gives a
comparison among the specific sorption volumes. The authors explain their
results with the help of the respective published data. Summing up: 1) The

Card 2/3

substitution of one part of the hydroxyl groups of the silica gel surface
by methyl groups leads to a reduction of the adsorptive power toward
substances in the vapor phase, and to an increase of the molecular fields
in the occupied monolayers. 2) To determine the specific surface
of the adsorbent with chemically non-homogeneous surface (on the basis
of the isotherms) by Brunauer, Emmett, and Teller for the adsorption iso-
therms. 3) It is necessary to select the adsorbates with special care.
They should be as little sensitive as possible to the chemical hetero-
geneity of the surface. L. M. Kurbatov is mentioned. There are 5 figures,
4 tables, and 22 references: 22 Soviet, 2 US, 1 British, 6 German and 1
Swiss.

ASSOCIATIONS: Institut Fizicheskoy Khimii Akademii Nauk SSSR (Institute
of Physical Chemistry of the Academy of Sciences USSR);
Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University named M. V. Lomonosov)

April 2, 1969

SUBMITTED:

Card 3/3

KISELEV, V.F.; KRASIL'NIKOV, K.G.

Effect of the nature of the silica surface on its adsorptive properties.
Probl. kin. i kat. 10:415-420 '60. (MIRA 14:5)

1. Fizicheskiy fakul'tet Moskovskogo gosudarstvennogo universiteta.
(Silica) (Adsorption)

KRASIL'NIKOV, K.G.; KISELEV, V.F.

Adsorptive properties of aluminosilicates and alumina gel. Probl.
kin. i kat. 10:421-425 '60. (MIRA 14:5)

1. Fizicheskiy fakul'tet Moskovskogo gosudarstvennogo universiteta.
(Aluminosilicates) (Alumina) (Adsorption)

67896

S/020/60/130/06/026/059

B004/BOC7

15.2110
~~5(4)~~

AUTHORS:

Kiselev, V. F., Krasil'nikov, K. G.,
Khodakov, G. S.

TITLE:

The Influence of the Aggregation of Quartz Particles During Grinding Upon Its Adsorptive Properties

PERIODICAL:

Doklady Akademii nauk SSSR, 1960, Vol 130, Nr 6, pp 1273 - 1276 (USSR)

ABSTRACT:

In reference 1 it was said that the specific surface of air-dried quartz decreases with an increase of the duration of grinding. This was explained by the aggregation of the quartz particles. The authors aimed at investigating this phenomenon more thoroughly and to find out whether its effects on the adsorption of nitrogen, and water differ. They maintain that this phenomenon is the cause of the considerable discrepancy in published data for adsorption values and adsorption energy of quartz. Two samples of highly dispersive quartz were investigated. Sample Kv-4 was obtained by grinding transparent crystalline quartz with an excess of water, sample Kv-4A by further grinding Kv-4 in air. On both samples, the adsorption of nitrogen and steam was measured (Table 1). As shown by

Card 1/3

67896

The Influence of the Aggregation of Quartz Particles S/020/60/130/06/026/059
During Grinding Upon Its Adsorptive Properties B004/B007

figure 1, the adsorption isothermal line of nitrogen on Kv-4A is lower than in the case of Kv-4 because of particle aggregation, whereas the adsorption isothermal line of steam is higher. Also figure 2 shows that the different kind of grinding the same quartz affects the adsorption of nitrogen and steam differently. This phenomenon has not yet been explained. It is presumed that relatively dense aggregates are formed, the inner surfaces of which are inaccessible to the nitrogen, whereas the adsorption of water is not impaired by these aggregations because of its dispersive (peptizing) properties. Such phenomena of aggregation were observed also in the case of other substances (corundum, calcite, silica gel) in dry grinding. The authors thank Academician P. A. Rebinder for his interest in this paper, and G. I. Aleksandrova for assisting in measurements. There are 2 figures, 1 table, and 21 references, 13 of which are Soviet.

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova
(Moscow State University imeni M. V. Lomonosov). Vsesoyuznyy
Card 2/3 nauchno-issledovatel'skiy institut novykh stroitel'nykh

67896

The Influence of the Aggregation of Quartz Particles S/020/60/130/06/026/059
During Grinding Upon Its Adsorptive Properties B004/B007

materialov (All-Union Scientific Research Institute for New
Building Materials)

PRESENTED: October 20, 1959 by P. A. Rebinder, Academician

SUBMITTED: October 13, 1959

4

Card 3/3

BONDARENKO, A.V.; KISELEV, V.F.; ^KBRASIL'NIKOV, K.G.

Composition of products of the thermal dehydration of silica and
properties of its dehydrated surface. Kin.i kat. 2 no.4:590-598
Jl-Ag '61. (MIRA 14:10)

1. Moskovskiy gosudarstvennyy universitet imeni M.V.Lomonosova,
fizicheskiy fakul'tet.
(Silica) (Dehydration (Chemistry))

GANICHENKO, L.G.; KISELEV, V.F.; KRASIL'NIKOV, K.G.; MURINA, V.V.

Effect of the nature of silica gel and quartz surfaces on their adsorption properties. Part 4: Adsorption and heat of adsorption of aliphatic alcohols on powdered silica gel. Zhur.fiz.khim. 35 no.8:1718-1726 Ag '61. (MIRA 14:8)

1. Moskovskiy gosudarstvennyy universitet imeni M.V. Lomonosova.

(Alcohols) (Adsorption) ✓

YEGOROV, M.M.; KISELEV, V.F.; KRASIL'NIKOV, K.G.

Effect of the nature of silica gel and quartz on their adsorptive capacities. Part 5: Structure of the surface of crystalline and amorphous modifications of silica. Zhur. fiz.khim. 35 no.9:2031-2038 '61. (MIRA 14:10)

1. Moskovskiy gosudarstvennyy universitet imeni M.V. Lomonosova.
(Silica) (Adsorption)

YEGOROV, M.M.; KISELEV, V.F.; KRASIL'NIKOV, K.G. (Moscow)

Effect of the nature of silica gel and quartz surfaces on their adsorption properties. Part 5: Adsorption of water vapor on the surface of crystalline and amorphous modifications of silica. Zhur.fiz.khim. 35 no.10:2234-2240 0 '61. (MIRA 14:11)

1. Moskovskiy gosudarstvennyy universitet imeni M.V.Lomonosova.
(Adsorption) (Silica)

BONDARENKO, A.V.; KISELEV, V.F.; KRASIL'NIKOV, K.G.

Thermal dehydration of silica and certain properties of its surface.
Dokl.AN SSSR 136 no.5:1133-1136 F '61. (MIRA 14:5)

1. Moskovskiy gos.universitet im. M.V.Lomonosova. Predstavleno akad.
M.M.Dubininym.

(Silica) (Dehydration) (Surface chemistry)

S/076/62/036/009/002/011
B101/B102

AUTHORS: Yegorov, M. M., Ignat'yeva, L. A., Kiselev, V. F., Krasil'ni-
kov, G. G., and Topchiyeva, K. V.

TITLE: Study of the surface properties of catalytic aluminum oxide

PERIODICAL: Zhurnal fizicheskoy khimii, v. 36, no. 9, 1962, 1882 - 1889

TEXT: The specific heat of wetting of commercial Al_2O_3 by water, methanol, ethanol, and n-heptane, and the content of structural water Al_2O_3 were measured, the phase composition of Al_2O_3 was determined by x-ray analysis, and the infrared spectrum of deuterated Al_2O_3 was taken. Whereas with n-heptane the heat of wetting is independent of the content of structural water in Al_2O_3 , it increases, in the case of water and alcohols, with increasing thermal dehydration of Al_2O_3 . Since, however, the specific surface of Al_2O_3 becomes smaller at high annealing temperatures, the heat of

Card 1/3

S/076/62/036/009/002/011
B101/B102

Study of the surface ...

wetting calculated per g of Al_2O_3 reaches a maximum for Al_2O_3 heated at 500°C. The curve for heat of wetting (Q , erg/cm²) versus structural water (μ mole/m²) shows the following sections: (1) Increase of Q after thermal treatment of Al_2O_3 at 20 - 150°C owing to removal of the adsorbed H_2O ; (2) unchanged Q at 170 - 230°C in spite of dehydration of the bayerit in the bulk of Al_2O_3 ; (3) Q increases at 290 - 500°C owing to dehydration of the Al_2O_3 surface; (4) sharp increase of Q between 500 and 700°C, although the content of structural water changes only little in this range owing to formation of γ - Al_2O_3 ; (5) increase of Q at 800-900°C owing to formation of K , δ , θ , and α - Al_2O_3 (corundum). The infrared spectrum of deuterated Al_2O_3 showed a broad 2630 cm⁻¹ band which disappeared at 400°C (interacting OD groups), a narrow band at 2755 cm⁻¹ (free, non-interacting OD groups), and a narrow 2710 cm⁻¹ band (weakly bound OD groups). For gibbsite, maximum hydration was calculated to be $\sim 22 \mu$ mole/m²; for the (0001) face of corundum, the hydration amounts to 12.7 μ mole/m². The coordination sphere of the Al

Study of the surface ...

S/076/62/036/009/002/011
B101/B102

surface atoms which is not fully occupied after the thermal dehydration is filled up by water or alcohols with formation of hydrate or alcoholates, respectively. The irreversible sorption of alcohols increases after thermal treatment of Al_2O_3 at high temperature. There are 4 figures and 2

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M. V. Lomonosova, Fizicheskii i khimicheskii fakul'tety (Moscow State University imeni M. V. Lomonosov, Physical and Chemical Departments)

SUBMITTED: November 1, 1960

KRASIL'NIKOV, K.G.

Sorption of water vapors on calcium hydrosilicates. Dokl. AN
SSSR 143 no.4:911-914 Ap '62. (MIRA 15:3)

1. Moskovskiy gosudarstvennyy universitet im. M.V.Lomonosova.
Predstavleno akademikom P.A.Rebinderom.
(Calcium silicates) (Water vapor) (Sorption)

KVLIVIDZE, V.I.; KRASIL'NIKOV, K.G.

State of water sorbed on calcium hydrosilicate studied by means of
nuclear magnetic resonance. Dokl.AN SSSR 145 no.6:1305-1307
Ag '62. (MIRA 15:8)

1. Moskovskiy gosudarstvennyy universitet im. M.V.Lomonosova.
Predstavleno akademikom P.A.Rebinderom.

(Water) (Sorption) (Nuclear magnetic resonance and
relaxation)

YEGOROV, M.M.; IGNAT'YEVA, L.A.; KISELEV, V.F.; KRASIL'NIKOV, K.G.;
TOPCHIYEVA, K.V.

Surface properties of catalytically active aluminum oxide.
Zhur. fiz. khim. 36 no.9:1882-1889 S '62. (MIRA 17:6)

1. Moskovskiy gosudarstvennyy universitet imeni Lomonosova,
fizicheskiy fakul'tet i khimicheskii fakul'tet.

KRASIL'NIKOV, K.G.

Calculation of the specific surface area of tobermorite from its
crystal lattice parameters and adsorption data. Dokl. AN SSSR
149 no.4:891-893 Ap '63. (MIRA 16:3)

1. Moskovskiy gosudarstvennyy universitet im. Lomonosova.
Predstavleno akademikom P.A.Rebinderom.
(Tobermorite) (Crystal lattices) (Adsorption)

GOROSHKO, O.A. [Horoshko, O.O.]; KRASIL'NIKOV, K.V. [Krasyl'nykov, K.V.]

Transverse vibrations of a string (cable) of variable
length. Dop. AN URSR no.3:319-322 '64. (MIRA 17:5)

1. Institut mekhaniki AN UkrSSR i Dnepropetrovskiy gosudarst-
vennyy universitet. Predstavleno akademikom AN UkrSSR G.N.
Savinym [Savin, H.M.].

RUBAN, P.I.; KRASIL'NIKOV, K.V.

One of the methods for the approximation of functions satisfying
Lipschitz's conditions by trigonometric polynomials. Izv.vys.
ucheb.zav.; mat. no.1:194-107 '60. (MIRA 13:6)

1. Dneprodzerzhinskiy vecherniy metallurgicheskiy institut imeni
Arsenicheva.

(Functions, Periodic)

RUBAN, P.I. (Dneprodzerzhinsk); KRASIL'NIKOV, K.V. (Dneprodzerzhinsk)

Approximation by trigonometric polynomials of functions of two variables satisfying Lipschitz' condition. Izv. vys. ucheb. zav.; mat. no.3:135-136 '63. (MIRA 16:4)

(Functions, Periodic) (Polynomials)

KRASIL'NIKOV, K.V., inzh.

Characteristics of the movement of flexible guides with pinched ends. Izv.vys.ucheb.zav.; gcr. zhur. 6 no. 12:141-145 '63.
(MIRA 17:5)

1. Dnepropetrovskiy gosudarstvennyy universitet. Rekomendovana kafedroy teoreticheskoy mekhaniki.

RUBAN, P.I. (Dneprodzerzhinsk); KRASIL'NIKOV, K.V. (Dneprozerzhinsk)

Use of trigonometric polynomials in the approximation of
even functions satisfying Lipshits' condition. Izv. vys.
ucheb. zav.; mat. no.2:136-138 '64. (MIRA 17:8)

3790L

S/137/62/000/005/117/150

A006/A101

128200

AUTHOR: Krasil'nikov, L. A.

TITLE: Some problems concerning methods of conducting stress-relaxation tests with wire

PERIODICAL: Referativnyy zhurnal, Metallurgiya, no. 5, 1962, 98, abstract 51601 ("Tr. Konferentsii po metizn. proiz-vu, 1959", Chelyabinsk, 178-183)

TEXT: The possibility was studied of using for thin wire the method of relaxation tests applied to a spring strip. According to this method, the relaxation stress is the difference between the initial and relieved stresses and is determined for both a straight and initially curved wire state. The method of determining stress-relaxation in a wire during plain bending should be used for a wire of ≤ 0.5 mm in diameter, as at a greater diameter, the measurement of the wire curvature becomes more difficult and, moreover, large-capacity furnaces or liquid baths are required. For more precise investigations of relaxation, a device was designed which makes it possible to obtain curves of relaxation processes in plain twisting. The twisting of the specimen under investigation on an operating model of the machine is brought about with the aid

Card 1/2

Some problems concerning methods...

S/137/62/000/005/117/150
A006/A101.

of a contour rotating due to interaction of magnetic fields inside two fixed coils; one of the wire-specimen ends is fixed on the axis of the movable contour and the other end in a fixed clamp. To establish a connection between the shear stress and voltage in the contour, the device is calibrated by small weights. The initial shear stress during the performance of relaxation tests is set by establishing a definite current value in the contour. To perform tests at elevated temperatures (100 - 600°C), the device is equipped with a furnace with automatic temperature control. 4

Ye. Assonova

[Abstracter's note: Complete translation]

Card 2/2

28(5)

AUTHORS:

Krasil'nikov, L. A., Karavayev, V. B.

SOV/32-25-7-32/50

TITLE:

Automatic Recording of the Relaxation of Tension in Wire
(Avtomaticheskaya registratsiya relaksatsii napryazheniy v
provoloke)

PERIODICAL:

Zavodskaya laboratoriya, 1959, Vol 25, Nr 7, pp 869-871 (USSR)

ABSTRACT:

In order to obtain correct data on the behavior of springs in use not only the mechanical properties have to be determined but also the resistance of the relaxation of tension (RT) and of elasticity at various states of tension (Refs 1-3). The resistance of (RT) is usually determined by measuring the rest deformation which appears gradually, caused by elastic deformation. A device was designed according to the principle of (Ref 4) by which the tension drop can be determined under the effect of a moment of torsion in a wire spring at 100-600° without measuring the deformation. The variation of the displacement module with temperature is determined together with the variation of tension. The testing conditions can be automatically controlled and the measuring results recorded by this device (Fig 1). The torsion of the sample is caused by the

Card 1/2

SOV/32-25-7-32/50

Automatic Recording of the Relaxation of Tension in Wire

interaction of two magnetic fields, by two stable coils and a frame rotating on its axis. The tension of relaxation is determined by means of a diagram (Fig 2) and the amperage in the frame which is set at a certain angle. During the test the sample is heated in a quartz tube furnace. The temperature is controlled and regulated by two electron potentiometers EPV-01. The measuring results obtained with wire samples of steel of the types 50KhFA and Kh18N9T (Fig 3) show that the primary displacement tension affect the relaxation stability at 200, 400, and 450° with regard to time. There are 3 figures and 5 Soviet references.

ASSOCIATION: Beloretskiy staleprovolochno-kanatnyy zavod (Beloretsk Steel Wire-Cable Factory)

Card 2/2

18.7190

770-1
007/1,000-4-21/85

AUTHORS: Golomanov, V. A., Krasnoludov, I. A. (Belobers)
TITLE: Improvement in Technological Process of Needle Wire
Production

PERIODICAL: Stal', 1960, Nr 2, pp 173-177 (USSR)

ABSTRACT: In order to improve mechanical properties as well as surface finish of needle wire made of U7A, 88A, and U10A-steel (carbon tool steels containing 0.7; 0.8; and 1.0% C, respectively) the authors worked on the improvement of the patenting process which has been used since 1953 at Beloretsk Steel Wire and Rope Plant (Beloretskiy staleprovolochnokanatnyy zavod). For that purpose the following factors were investigated: (1) Effect of preliminary heat treatment and total deformation on recrystallization annealing: Industrial and laboratory tests showed that wire subjected to recrystallization annealing after different types of

Card 1/7